

NAME:		ROLL NO:		TIME:	50: MIN
SUB:	Chemistry	Date:		CLASS:	1 st Year
SYLLABUS	Chp.1 (Periodic Table and Periodic Properties)	T. MARKS	30	OBT. M	

Q. 1 Choose the Correct Answer. Cutting, Erasing, Double Selection is Strictly Prohibited. (1x6=6)

1. Which of the following elements is not a member of the chalcogen family?

- (A) Oxygen (O) (B) Sulfur (S) (C) Selenium (Se) (D) Bromine (Br)

2. Non-metals are present in which block of periodic table?

- (A) S-block (B) P-block (C) D-block (D) F-block

3. Which of the following has the highest electron affinity?

- (A) Fluorine (B) Chlorine (C) Bromine (D) Iodine

4. Among the following members of group 14 which is a non-metals?

- (A) C (B) Si (C) Ge (D) Sn

5. The oxides of electropositive elements are:

- (A) Acidic (B) Basic (C) Amphoteric (D) Neutral

6. Oxidation number of AL in Al_2O_3 is:

- (A) +2 (B) +3 (C) +4 (D) +5

PART-I

Q. 2 Write Short Answers of the Following Questions (2x8=16)

- Why are elements in group 1 and 2 known as s-block elements?
- How does electron affinity varies in a group of the periodic table?
- Using your knowledge of period 3 elements, predict and explain the relative sizes of the atomic radii of lithium and fluorine.
- Explain with reasoning following facts about ionization energy 1st ionization energy of Boron is lesser than Beryllium.
- What is the nature of oxides and hydroxides of Na and Mg?
- ZnO reacts with HCL to give $ZnCl_2$ and with NaOH to give Na_2ZnO_2 . Give equations and also predict the type of the oxide?
- Discuss any two factors affecting electronegativity.
- What is the role of Shielding effect on ionization energy?

PART-II

Write Answers of the Following Questions in Detail.

Q. 3 (a) Define 1st and 2nd ionization energy with one example each. Also discuss periodic trends in ionization energy. 4

(b) Write down the reactions of Na and Mg with water, oxygen and chlorine (equations only). 4



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1ST YEAR REVISION & TEST SESSION



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Answer Key

1	D	2	B	3	B	4	A
5	B	6	B				