

NAME:		ROLL NO:		TIME:	50: MIN
SUB:	Chemistry	Date:		CLASS:	1 st Year
SYLLABUS	Chp. 13 (Halogens)	T. MARKS	30	OBT. M	

Q. 1 Choose the Correct Answer. Cutting, Erasing, Double Selection is Strictly Prohibited. (1x6=6)

- Which halogen molecule has the strongest bond?
(A) F₂ (B) Br₂ (C) I₂ (D) Cl₂
- Which one the following halogen molecule has strongest oxidizing power?
(A) Br₂ (B) F₂ (C) I₂ (D) Cl₂
- The decreasing thermal stability of the halogens down the group is primarily due to the:
(A) Increasing electronegativity of the atoms
(B) Decreasing bond length between the halogen atoms
(C) Increasing atomic radius leading to a weaker covalent bond
(D) Increasing strength of Van der Waals forces
- How does the acidic strength of hydrogen change as you move down the group?
(A) It remains constant (B) It decreases from HF to HI
(C) It increase from HF to HI (D) It fluctuates erratically
- Why is fluorine the most reactive halogen?
(A) Bond length in the halogen molecule (B) Bond strength in the halogen molecule
(C) Electronegativity of the halogen (D) Number of electrons in the halogen molecule
- When aqueous silver nitrate is added to a solution containing bromide ions, a cream precipitate forms. What is the solubility of this precipitate in ammonia solution?
(A) Soluble in dilute ammonia solution (B) Partially soluble in dilute ammonia solution
(C) Insoluble in dilute ammonia solution (D) Soluble only upon heating with ammonia

PART-I

Q. 2 Write Short Answers of the Following Questions (2x8=16)

- Why chlorine is more volatile than bromine and iodine?
- HF is the most thermally stable hydrogen halide. Give reasons.
- Write down the equation for the reaction of KI with Ag⁺ followed by NH₃. What would you observe at the completion of this reaction?
- HI acts as strong reducing agent. Explain it with two chemical reactions.
- Which halogen is used as an antiseptic? How does it work?
- Why HF is weaker acid than HCl?
- Describe a simple chemical test that could be used to distinguish between aqueous solutions of potassium bromide and potassium iodide. Include the reagents and expected observations.
- Explain the chemical principles behind the use of chlorine as a disinfectant in water purification. Include relevant chemical equations in your explanation.

PART-II

Write Answers of the Following Questions in Detail.

- Q. 3 (a) Discuss reactions of halides with aqueous silver ion followed by aqueous ammonia. 4
- (b) Describe the reactions that occur when chlorine is bubble through 4
- (i) Cold and (ii) Hot, aqueous sodium hydroxide (NaOH).



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1ST YEAR REVISION & TEST SESSION



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Answer Key

1	D	2	B	3	C	4	C
5	C	6	B				