

NAME:		ROLL NO:		TIME:	15: MIN
SUB:	Chemistry	Date:		CLASS:	1 st Year
SYLLABUS	Chp. 5 (States and Phases of Matter)	T. MARKS	11	OBT. M	

Q. 1 Choose the Correct Answer. Cutting, Erasing, Double Selection is Strictly Prohibited. (1x11=11)

- London dispersion forces are the only forces present among:**
 - Molecules of water in liquid state
 - Atoms of helium in gaseous state at high temperature
 - Molecules of solid iodine
 - Molecules of hydrogen chloride gas
- When the vapour pressure of a liquid equals the external pressure, what phenomenon occurs?**
 - Sublimation
 - Condensation
 - Boiling
 - Freezing
- When water freezes at 0°C, its density decreases due to:**
 - Cubic structure of ice
 - Empty spaces present in the structure of ice
 - Decreases in volume
 - Decreases in viscosity
- Which of the following is the correct sequence of decreasing ΔH_v values of substances mentioned?**
 - $H_2O > NH_3 > F_2$
 - $F_2 > NH_3 > H_2O$
 - $NH_3 > H_2O > F_2$
 - $H_2O > F_2 > NH_3$
- Surface tension of a liquid is due to:**
 - Inward pull of surface molecules
 - Upward pull from the surface
 - Collision of molecules
 - Repulsive forces
- Which change of state involves overcoming only London dispersion forces?**
 - Melting of ice ($H_2O_{(s)} \rightarrow H_2O_{(l)}$)
 - Boiling of ethanol ($CH_3CH_2OH_{(l)} \rightarrow CH_3CH_2OH_{(g)}$)
 - Sublimation of iodine ($I_{(s)} \rightarrow I_{(g)}$)
 - Dissolving sodium chloride in water
- In which of the following substances are London dispersion forces the only significant intermolecular forces present?**
 - Ammonia (NH_3)
 - Water (H_2O)
 - Methane (CH_4)
 - Hydrogen fluoride (HF)
- Which of the following is a characteristic property of crystalline solids?**
 - They have a range of melting points.
 - They are isotropic
 - They have a definite and sharp melting point
 - They lack a regular arrangement of particles.
- Which of the following liquids would you expect to have the highest viscosity at a given temperature?**
 - Water (H_2O)
 - Ethanol (CH_3CH_2OH)
 - Diethyl ether ($CH_3CH_2OCH_2CH_3$)
 - Glycerol ($CH_2(OH)CH(OH)CH_2(OH)$)
- Which type of intermolecular force is present in all types of molecule of their polarity?**
 - Dipole-dipole forces
 - Hydrogen bonds
 - London dispersion forces
 - Ion-dipole forces
- Liquid crystals exhibit properties:**
 - Only like solids.
 - Only like liquids.
 - Between solids and liquids.
 - Unlike solids or liquids.



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1ST YEAR SOLVED EXERCISE MCQs



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Answer Key

1	B	2	C	3	B	4	A
5	A	6	C	7	C	8	C
9	D	10	C	11	C		