

NAME:		ROLL NO:		TIME:	25: MIN
SUB:	Chemistry	Date:		CLASS:	1 st Year
SYLLABUS	Chp. 13 (Halogens)	T. MARKS	20	OBT. M	

Q. 1 Choose the Correct Answer. Cutting, Erasing, Double Selection is Strictly Prohibited. (1x20=20)

- Which halogen molecule has the strongest bond?**
 (A) F_2 (B) Br_2 (C) I_2 (D) Cl_2
- The volatility of the halogens (Group 17) generally as you move down the group (from Fluorine to Iodine).**
 (A) Increases (B) Decreases
 (C) Remains the same (D) Fluctuates unpredictably
- Which one of the following halogen molecules has strongest oxidizing power?**
 (A) Br_2 (B) F_2 (C) I_2 (D) Cl_2
- The decreasing thermal stability of the halogen down the group is primarily due to the:**
 (A) Increasing electronegativity of the atoms.
 (B) Decreasing bond length between the halogen atoms.
 (C) Increasing atomic radius, leading to a weaker covalent bond.
 (D) Increasing strength of van der Waals forces
- Which one of the following halides has strongest reducing power?**
 (A) F^- (B) Cl^- (C) Br^- (D) I^-
- Which statement about the reaction between halogens and hydrogen is correct?**
 (A) Iodine reacts most vigorously with hydrogen
 (B) Chlorine and hydrogen explode in darkness
 (C) Fluorine combines explosively with hydrogen even in cold and dark conditions
 (D) Bromine and hydrogen do not react at all
- How does the acidic strength of hydrogen halides change as you move down the group?**
 (A) It remains constant. (B) It decreases from HF to HI.
 (C) It increases from HF to HI (D) It fluctuates erratically
- Why is fluorine the most reactive halogen?**
 (A) Bond length in the halogen molecule (B) Bond strength in the halogen molecule
 (C) Electronegativity of the halogen (D) Number of electrons in the halogen molecule
- When aqueous silver nitrate is added to a solution containing bromide ions, a cream precipitate forms. What is the solubility of this precipitate in ammonia solution?**
 (A) Soluble in dilute ammonia solution. (B) Partially soluble in dilute ammonia solution.
 (C) Insoluble in dilute ammonia solution. (D) Soluble only upon heating with ammonia.
- Concentrated sulfuric acid is added to solid sodium chloride. What is the initial observation?**
 (A) Reddish-brown fumes are evolved.
 (B) A purple vapor is evolved.
 (C) Steamy white fumes of hydrogen chloride are evolved.
 (D) A black solid is formed.

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11. Which halogen has greenish yellow colour?

- (A) F₂ (B) Cl₂ (C) Br₂ (D) I₂

12. Which halogen exists in liquid state?

- (A) F₂ (B) Cl₂ (C) Br₂ (D) I₂

13. Which of the following is a strongest reducing agent?

- (A) F⁻ (B) Br⁻ (C) Cl⁻ (D) I⁻

14. Oxidizing power of which halogen is highest?

- (A) F₂ (B) Cl₂ (C) Br₂ (D) I₂

15. Which is strongest acid?

- (A) HF (B) HCl (C) HI (D) HBr

16. Highest electron affinity is shown by?

- (A) F (B) I (C) Br (D) Cl

17. Which is the strongest reducing agent?

- (A) HF (B) HCl (C) HI (D) HBr

18. Substance boiling at higher temperature among following is?

- (A) HI (B) HF (C) HCl (D) HBr

19. Group VII-A elements are generally called:

- (A) Halogens (B) Noble gases (C) Inert gases (D) Metalliods

20. The radioactive element in halogen group is:

- (A) Radon (B) Radium (C) Astatine (D) Bromine



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Answer Key

1	C	2	B	3	B	4	C
5	B	6	C	7	C	8	B
9	B	10	C	11	B	12	C
13	D	14	A	15	D	16	D
17	C	18	B	19	A	20	C